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Note: Not all of the reactions will occur. For those that do not, write no reaction.

1. A piece of copper is dropped into a container of water. No reaction
2. Liquid bromine is added to a container of sodium iodide crystals. $\text{Br (l)} + 2\text{NaI(s)} \rightarrow 2\text{NaBr(s)} + \text{I}_2\text{(s)}$
3. An aluminum strip is immersed in a solution of silver nitrate.

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Pg. 50, Practice-Pg. 51 • Acid Base Neutralization Reactions- Pg. 53, Practice-Pg.54 oWeak acids and bases do not break apart and are in net ionic equation. • Synthesis and Decomposition Reactions - Pgs. 42-43, Practice- Pg. 43 • Single Replacement Reactions - Pg. 45, Practice-Pg. 46.

Quiz on all Chemical Reactions - Max Study

The Ultimate Chemical Equations Handbook SnCO tin(II) carbonate sodium hydrogen carbonate NaHCO 1. 2. 3. 4. 5. 6. 8. 9. 10. 11. vanadium(V) oxide H₂O dihydrogen monoxide (NH) CO ammonium oxalate 4 224 polonium(VI) thiocyanate Po(SCN)₆ tetraphosphorus decaoxide P₄O₁₀ 12. 13. 15. 16. 17. 18. 19. manganese(VII) oxide copper(II) dihydrogen phosphate

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Inc. Correlating questions (ii) from:
Rodney Schmitz (2007) Moore County
Schools, NC Reaction Prediction - 5 For
each of the following eight reactions, in
part (i) write a BALANCED equation and

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In part (ii) answer the question about the reaction.

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The Ultimate Chemical Equations Handbook Determine the oxidation number of the underlined element:
 KMnO_4 Since K is an alkali metal, its charge must be $1+$ Oxygen is $2-$ but there are four of them, therefore, 4 times $2-$ equals $8-$ If $1+$ and $8-$ are added Ultimate Chemical Equation Answers

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The Ultimate Chemical Equations Handbook Determine the oxidation number of the underlined element: KMnO_4 Since K is an alkali metal, its charge must be 1+ Oxygen is 2— but there are four of them, therefore, 4 times 2— equals 8— If 1+ and 8— are added The Ultimate Chemical Equations Handbook Answers Chapter 10

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