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Chapter 6 Solutions Engineering and Chemical Thermodynamics Wyatt Tenhaeff Milo Koretsky Department of Chemical Engineering Oregon State University 2 6.1 (a) The Clausius-Clapeyron equation: $dP_i^{\text{sat}}/P_i^{\text{sat}} = \frac{h_i^{\text{vap}}}{RT^2} dT$ or $\ln P_i^{\text{sat}} = \frac{h_i^{\text{vap}}}{R} \left(\frac{1}{T} - \frac{1}{T_0} \right) + \ln P_i^{\text{sat}}(T_0)$ where $P_i^{\text{sat}}(T_0) = 101 \text{ kPa}$ at $T_0 = 373 \text{ K}$. So $P_i^{\text{sat}} = 101 \dots$

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Chapter 6 2 If the fluid is in thermodynamic equilibrium any thermodynamic variable for a pure substance, like pure water, can be written in terms of any two other thermodynamic variables, i.e. $p = p(p, T)$ (6.1.1) where the functional relationship in depends on the substance.

Chapter 6 Thermodynamics and the Equations of Motion

Chapter 6: Solution Thermodynamics and Principles of Phase Equilibria In all the preceding chapters we have focused primarily on thermodynamic systems comprising pure substances. However, in all of nature, mixtures are ubiquitous.

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